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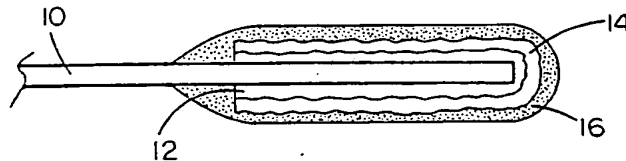
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(54) **Solid state electrode.**

(57) The invention described herein is a sterilizable reference electrode suitable for use in medical applications. The electrode is formed by dip-coating a metal/metal halide electrode in a saturated aqueous solution of a soluble salt of the halide. The electrode is then dried and dipped in a nonaqueous solution of a hydrophobic membrane material to produce a layer over the salt layer which is permeable to ionic current-carrying species in a nonspecific manner. The resulting electrode can then be inserted into a lumen of a catheter and sterilized using ethylene oxide gas or other standard methods. Also described is a sterilizable ion-specific electrode which is suitable for medical applications and which is similar to the reference electrode except for the replacement of the nonspecific membrane with a species-selective membrane.



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BACKGROUND OF THE INVENTION

The invention relates to electrodes and more specifically to miniaturized reference electrodes for in vivo medical applications.

05 Those concerned with monitoring various ionic or gaseous species in blood typically need to be able to use a miniaturized electrode in vivo to obtain continuous real-time indications of the factors to be monitored. One method of obtaining the necessary information is to
10 use miniaturized electrochemical sensors consisting of a reference electrode and a species-selective electrode in contact with the blood flow of a patient. The most convenient configuration for the electrode pair is to be mounted in a lumen of a catheter which may then be in-
15 serted into the bloodstream of the patient through a vein or an artery.

 In the past, each individual electrode of a pair of electrodes typically required an aqueous phase within its lumen. This aqueous phase was necessary to
20 make contact between the electrode and the external environment, as well as to provide a reference solution having a fixed concentration of a particular ion required to establish a stable potential for the electrodes. This created a problem in medical applications because it made
25 it difficult to sterilize the electrical sensors by means of the standard procedures using ethylene-oxide gas. When ethylene-oxide is used in the sterilization procedure for any object containing water, a reaction can occur between the ethylene-oxide gas and the water to
30 form ethylene glycol, which is quite toxic. Obviously, this is highly undesirable for medical applications. Furthermore, the ethylene glycol will change the nature

of the electrode reference solution. Thus, a need existed to develop an electrode which is completely anhydrous, and therefore sterilizable, and yet would function satisfactorily.

05 An example of the prior art is described in U.S. Patent No. 3,856,649 to Genshaw et al. In this patent, a solid-state electrode is described for determining ion concentrations in an aqueous solution. The electrode includes an electrically conductive inner
10 element with a salt disposed on a surface portion thereof having a cation and an anion. The cation is identical to at least a portion of the inner electrode material. A solid hydrophilic layer is in intimate contact with the salt and includes a water-soluble salt of the anion. It
15 is important to note that in the device described in the Genshaw reference, the soluble salt must be mixed with the hydrophilic layer which may or may not be dried after application, but which must be wet in order for the electrode to operate. This is undesirable because such an
20 electrode cannot be sterilized with ethylene oxide while wet, and if dried for sterilization, requires a "wet-up" time before use. Furthermore, the electrode described by Genshaw et al. includes a polyvinyl-chloride hydrophobic layer which the present inventors have found to be highly
25 unsatisfactory for the fabrication of reference electrodes. Plasticized polyvinyl-chloride is unsatisfactory because it appears to be inherently selective to certain ions, in particular K^+ .

30 Another ion-sensing device for medical applications is described by Band, D.M. and Treasure, T. J. Physiol. (London), 266, 12, 1977 and reprinted in Ion-Selective Electrode Methodology, Vol. II, 1979, 58.

This is a potassium-sensing catheter for in vivo use, and consists of a potassium-sensing element and a reference electrode which are both inserted in separate lumens of a single catheter. The reference electrode consists of a
05 silver/silver chloride wire that is immersed in a saturated solution of potassium chloride which fills the lumen and which must be periodically flushed and replaced. Obviously, this is inconvenient for both medical personnel and patients and is an undesirable feature
10 because the accuracy of electrical readings using the electrode will suffer if the potassium-chloride solution is not periodically refreshed. This type of arrangement is typical of most in vivo chemical sensing devices. In view of the limitation of these types of prior art
15 devices, a need existed to provide a chemical sensor for in vivo applications which is completely dry and capable of being sterilized with ethylene-oxide gas.

SUMMARY OF THE INVENTION

An object of the subject invention, therefore,
20 was to provide a chemical sensor for in vivo applications which is completely dry and capable of being sterilized with ethylene-oxide gas.

Another object of the subject invention is to provide a device which can be easily fabricated.

25 Still another object of the subject invention is to provide a reference electrode which can be used as part of an electrochemical sensor for in vivo medical applications and which does not require intermittent flushing or other forms of generally continuous maintenance to insure the accuracy of electrical readings
30 using the electrode.

Yet another object of the subject invention is to provide a reference electrode which is completely dry and does not require a wetting period in order for the electrode to function.

05 Another object of the subject invention is to provide a truly stable reference electrode which is not sensitive to chemical changes in the external environment.

10 Another object of the subject invention is to provide an electrode sufficiently small to be inserted into a lumen of a catheter.

 Another object of the subject invention is to provide a class of electrodes for selectively sensing specific ionic species.

15 Another object of the subject invention is to provide a membrane for a reference electrode which is ionically conductive but nonselective.

20 These and other objects and advantages of the invention, as well as the details of an illustrative embodiment, will be more fully understood from the following description and the drawings.

 The invention can be summarized as an electrode for use in miniaturized sensor applications. The sensor includes an electrically conductive inner element and a
25 first salt layer on the surface of the electrically conductive layer having a cation in common with some portion of the electrically conductive inner element. The first salt layer also has an anion. The invention also includes a second salt layer coating the first salt layer,
30 the second salt layer having an anion in common with the first salt layer, the second salt layer being totally anhydrous. In various embodiments of the invention, the

electrode can be used as either a reference or sensing electrode. These embodiments are discussed in greater detail below.

BRIEF DESCRIPTION OF THE DRAWINGS

05 FIG. 1 is a cross-sectional view of a reference electrode illustrating the various layers of material used in a reference electrode in one embodiment of the subject invention;

10 FIG. 2 is a view similar to FIG. 1, illustrating one embodiment of the subject invention in which a reference electrode and a sensing electrode are inserted in separate lumens of a catheter;

15 FIG. 3 is a graph illustrating the insensitivity of the response of a reference electrode of the subject invention to changes in potassium concentration and a comparison of the insensitivity of the subject invention electrode with that of prior-art electrodes; and

20 FIG. 4 is a graph illustrating the K^+ response of a reference electrode and K^+ sensing electrode, mounted in a catheter, to changes in the K^+ concentration.

DETAILED DESCRIPTION OF THE INVENTION

25 Referring now to FIG. 1 illustrating the various layers used in an electrode described by the subject invention, it is important to note that the invention comprehends the use of a dry electrode for both reference and sensing applications. The electrode can be

either a reference or sensing electrode depending on the type of membrane applied to the outer surface of the electrode. As can be seen in the figure, a wire-type structure 10 is used as the conductive element of the electrode 11. Typically, the conductive element will be formed from silver or a silver alloy. Silver is preferred because it is generally chemically inert since it is a noble element. However, other materials can be used for the inner conductive element for in vivo applications provided that they have the following characteristics: first, they must be nonreactive with body fluids; second, they must be capable of forming an insoluble salt; and third, they must be electrically conductive. Another desirable feature of the material used to form the inner conductive element is that it be capable of forming a reversible electrode with its cation. Other materials that could be used as the inner conductive element are mercury, thallium, and lead. However, each of these materials has serious drawbacks as compared to silver, most notably the fact that the latter two materials are generally more reactive and therefore more likely to become highly toxic in physiological applications, and mercury is less convenient to work with because it is a liquid.

The electrode 11 of the subject invention is formed by applying a first layer 12 composed of salt to the outside surface 13 of the inner conductive element 10 (FIG. 1). This layer is applied by electrochemically oxidizing element 10 in a solution of a soluble salt of an anion which forms a precipitate with the cation of the metal conductor of element 10. In the preferred embodiment, the first salt layer 12, therefore, has a cation in common with the metal of the inner conductive layer. Thus, when the inner conductive element 10 is

composed of silver, the first salt layer 12 will form an insoluble silver salt. The first salt layer is required to be a relatively insoluble salt so that the concentration of the cation of the first salt layer 12 will be fixed as a result of an equilibrium with a second salt layer 14 discussed below and yet will not contribute appreciably to the ion current passing through a permeable membrane 16. In the preferred embodiment, the first salt layer is composed of silver chloride formed by the electrochemical oxidation with the silver wire inner conductive element 10 in a dilute solution of HCl. In general, the thickness of the first salt layer should be approximately one third the radius of the element 10. This is to insure that sufficient silver chloride remains after subsequent fabrication procedures discussed below, while at the same time leaving sufficient silver wire to retain mechanical strength or integrity.

In the preferred embodiment, application of the first salt layer upon element 10 can be enhanced by using the procedures discussed below. First, a 0.020" diameter silver wire used as element 10 is annealed prior to the application of the first salt layer. Second, the annealing step is followed by brief etching of the surface of the silver wire in nitric acid for cleaning purposes. Third, after etching, the surface of the wire is further treated by soaking in ammonium hydroxide to remove soluble impurities including silver-oxide. Fourth, the wire is then electrochemically oxidized to apply the first salt layer. The solution of HCl in which the wire is oxidized can vary. Generally, however, the concentration of the solution should be in the range of 0.01 to 0.5 molal. Typically, the current applied to the electrode is in the range of 1 to 100 mA/cm². In the

preferred embodiment, a first cycle of current of 50 mA/cm² is applied to element 10 for one minute. The polarity of the current is then reversed for one minute to electrochemically reduce a portion of the silver chloride to cause particles of silver to be present in the silver-chloride layer. This provides an integral mixture of silver particles in the silver-chloride salt layer 12. And finally, the polarity of the current is again reversed, the current density is reduced to about 10 mA/cm², and is applied for approximately 10 minutes.

As can be seen in FIG. 1, a second salt layer 14 is formed upon the first salt layer 12. The purpose of the second salt layer is to provide a fixed concentration of the anion of the first salt layer 12 so as to define the equilibrium concentration of the cation of the first salt layer 12. Another purpose of the second salt layer is to provide majority ionic current carriers through an outer membrane of the electrode as discussed below. A very important consideration in the selection of a material for the second salt layer 14 is that the ionic mobilities of the cation and anion be as nearly equal as possible.

The subject invention comprehends that the second salt layer 14 is applied by dipping the previously described silver/silver chloride covered element 10 into a saturated solution of the salt used to form the second layer 14 of salt. It is advantageous for the salt solution to be at a superambient (elevated) temperature during dipping in order to achieve a fine coating of second-layer salt particles on the surface of the first layer as well as to facilitate the evaporation of water from the electrode surface after dipping. It is desirable to have a fine particle size in order to reduce the overall thickness of the electrode while still

maintaining the electrochemical equilibrium as discussed above. This is important to reduce the bulk of the electrode as much as possible.

05 In the preferred embodiment, the second salt layer 14 is applied upon layer 12 by dipping the element 10 in a saturated solution of potassium chloride. Typically the element and first salt layer is dipped three times in a solution maintained at 80°C. The structure is allowed to dry between dippings. This occurs almost immediately after dipping. The number of dippings is not critical. In some embodiments, a single dipping may be sufficient. Typically, the thickness of the second salt layer is considered to be sufficient when a uniform white layer of potassium chloride is visible on the surface of the electrode.

15 Although in the preferred embodiment potassium chloride is used to form the second salt layer, other materials can be used. For example KNO_3 , NH_4NO_3 , or RbCl may be used as the second salt layer 14, although in the cases of KNO_3 and NH_4NO_3 , a second soluble salt containing the chloride anion would need to be present at some small but finite concentration. The main requirements of the material chosen to form the second salt layer 14 are: first, the cation and anion should have nearly the same ionic mobility; second, the salt should be readily soluble in water (water solubility is important to be able to provide a high concentration of the majority current carriers); and third, it should contain the anion of the first salt layer. This second salt layer 14 can be composed of a mixture of two or more salts. When used to form a sensing electrode, the composition of this second salt layer 14 may require

the inclusion of other ionic species with particular properties (e.g., identity with the ion being sensed). This will be discussed in greater detail below.

When the subject invention is used as a reference electrode, a hydrophobic outer layer 16 is formed upon layer 14. The outer layer is a membrane which is not selective with respect to ion transport by a particular ion, thus allowing the structure to operate as a reference electrode. In the preferred embodiment, the substance used as the membrane is plasticized or unplasticized cellulose acetate butyrate. The membrane is applied by dipping the structure containing the first and second salt layers into a solution of the membrane material in a nonaqueous solvent such as tetrahydrofuran. The thickness of the membrane layer ranges between 0.05 and 0.8 mm. In the preferred embodiment, the thickness is not critical provided that the membrane substance completely covers the internal salt layers. It is also necessary that the membrane be sufficiently thick to avoid the presence of pinholes and to provide sufficient mechanical stability. However, if the membrane becomes too thick, the electrode resistance and its sensitivity to electrical noise will become excessively high.

It should be noted that other materials may be used as the membrane material when the subject invention is used as a reference electrode. Other materials include any plasticized or unplasticized polymeric material taken from the group consisting of cellophane, collodion, or cellulosic derivatives such as cellulose acetate butyrate, ethyl cellulose, cellulose acetate, cyanoethylated cellulose, cellulose propionate, or cellulose tridecanoate. However, cellulose acetate butyrate or ethyl cellulose plasticized with didecylphthalate is preferred.

Referring now to FIG. 3 which shows a plot of the response 32 of a reference electrode formed using the techniques described herein to changes in the concentration of potassium chloride in an aqueous solution relative to an electrode of fixed potential. As can be seen in the figure, there is no response to changes in the concentration of the solution for the subject electrode. Also shown in FIG. 3, is a plot of the response 34 of an electrode made using a polyvinyl chloride membrane of the type described in U.S. patent No. 3,856,649 to Genshaw et al. As can be seen from the figure, the electrode made using a polyvinyl chloride membrane shows a substantial response to changes in the potassium-ion concentration.

The electrode 11 of the subject invention can be used as a sensing electrode as well as a reference electrode provided that a different membrane is used. In a sensing application, the outer membrane must be selective to the particular ion to be monitored. For instance the subject electrode can be used to monitor a multitude of chemical species such as the ions of potassium, sodium, ammonium, calcium, magnesium, or any other of the various ionic species.

Two key features of the invention which are discussed in greater detail below are that the electrode 11 is basically anhydrous, and does not have to be intentionally hydrated to function, and that it produces a thermodynamically well-defined potential. The electrode has a thermodynamically well-defined potential because all phases involved in the electrode reaction are clearly defined and are mutually at equilibrium.

When the electrode is used to sense potassium, the outer layer 16 could be formed from silicone rubber or plasticized polyvinyl chloride, polyhydroxyethyl methacrylate or a number of other similar hydrophobic polymeric materials. Any of the polymers used may be plasti-

cized using compatible plasticizers, such as didecylphthalate in the case of polyvinyl chloride, or any of the class of phthalic acid derivatives, citric acid derivatives, adipic acid derivatives, or sebacic acid derivatives, among others. In order to achieve selectivity for potassium ions, valinomycin, crown ethers or some similar specific complexing agent is added to the solution of the polymeric membrane material before the electrode is coated. Depending on the particular ion to be monitored, a variety of materials may be used for the ion-selective membrane. For example, in one embodiment of the subject invention, a mixture may be formed using polyvinyl chloride, didecylphthalate, and a complexing agent selective to said specific ionic species. In the preferred embodiment, a 47.5% polyvinyl chloride, 47.5% didecylphthalate, 5% Valinomycin membrane is used to selectively measure K^+ concentrations.

Referring now to FIG. 4 which shows a plot of the response characteristic of an electrode pair consisting of a potassium-selective electrode 30 relative to a reference electrode 28, both constructed as described in the subject application, and mounted in a catheter as shown in FIG. 2. FIG. 4 demonstrates that the sensitivity of the electrode pair to changes in the potassium concentration is nearly equal to the theoretical value, being 55 mV/pK⁺ as compared to the theoretical value of 59 mV/pK⁺.

In a similar manner, when the electrode 11 is used to sense other ions, an appropriate complexing agent specific for that particular ion is added to the polymeric solution in place of valinomycin or crown ether.

It is important to note that in the subject invention, the electrode is completely anhydrous when formed. This distinguishes it from other devices in the field which typically require aqueous solutions or hydro-

phylic layers which absorb and retain water. Although the detailed mechanism of operation of the subject electrode is not yet completely understood, it is believed that trace amounts of water are absorbed through the membrane during operation. This produces a saturated solution of the second layer salt; the solution composition will remain fixed as long as some solid from the second salt layer 14 remains.

Another important feature of this type of sensing electrode compared with prior art coated-wire electrodes is that all phases and interfaces of sensors fabricated in accordance with the techniques described herein are thermodynamically well-defined so that the potential of the electrode has thermodynamic significance. The potential of the prior-art coated wire sensor is determined by capacitive effects and is therefore less stable and reproducible.

Referring again to FIG. 2, in one embodiment of the invention, a reference electrode and an ion-selective sensing electrode are inserted in separate lumens of a single catheter. A catheter 18 has a first and second lumen 20, 22. Each lumen has an orifice 24, 26 to expose a portion of each lumen to the environment. A reference electrode 28 is inserted in one catheter of lumen 20, and a sensing electrode 30 is inserted in the other lumen. After each of the electrodes are inserted, each orifice is filled with a membrane material identical to the membrane material used for the specific electrode within the lumen. The solvent for the membrane solution is chosen such that it will partially dissolve the material of the catheter during application thereby forming a sealing bond between the membrane and the catheter.

It is also possible to insert both electrodes into the same lumen if care is taken to insulate the two electrodes and electrical conductors from one another. Thus, it is possible to form a sensing device which is specific for a given ionic species, can be sterilized using ethylene oxide or gamma sterilization, and is suitable for in vivo applications. In addition, the electrode sensing devices formed using the techniques of the subject invention are simple to prepare and convenient to use. Since sensors formed using the subject invention can be easily sterilized without fear of producing harmful toxins, such as ethylene glycol, such sensors are particularly appropriate for medical applications.

A major advantage of the subject invention is that electrodes formed using the techniques described herein can be sterilized for medical purposes by various techniques including the use of radiation, gas, ultraviolet light and steam.

Another advantage of the subject invention is that since sensors made using this technique contain only solid salt phases and polymer membranes, the technique lends itself particularly to the fabrication of extremely small sensing devices. For example, such sensor can be incorporated into electronic devices such as field effect transistors. It is further envisioned that sensors fabricated using the techniques described herein could even be sufficiently miniaturized to be placed within a hypodermic needle or medical cannula.

Although the invention has been described and illustrated in detail, it is to be clearly understood that the same is by way of illustration and example only, and is not to be taken by way of limitation; the spirit and scope of this invention being limited only by the terms of the appended claims.

WE CLAIM:

1. An electrode for use in miniaturized sensor applications, comprising:

an electrically conductive element;

05 a first layer of salt on the surface of said electrically conductive element having a cation in common with some portion of said electrically conductive element, said first layer also having an anion; and

10 a second layer of salt providing a coating on said first layer, said second layer establishing an anion in common with said first salt layer, and said second salt layer being totally anhydrous.

2. An electrode as recited in Claim 1, wherein:

05 said second salt layer provides an anhydrous coating formed by dipping said conductive inner element including said first salt layer into a saturated solution of said salt or salts of which said second layer is to be formed to cause an anhydrous coating of said salts to be formed on said first salt layer.

3. An electrode as recited in Claim 2, wherein:

05 said saturated solution of said second salt layer is at a superambient temperature during said dipping.

4. An electrode as recited in Claim 3, wherein:

said second salt layer includes a majority component selected from the group consisting of KCl, KNO₃, NH₄NO₃, and RbCl.

5. An electrode as recited in Claim 4, wherein:

said second salt layer includes a minority component for establishing said anion in common with said first salt layer and a cation.

6. An electrode as recited in Claim 3, wherein:

said second salt layer is formed of anhydrous KCl.

7. An electrode as recited in Claim 1, wherein said electrode is a reference electrode further comprising:

a reference membrane surrounding said second salt, said membrane being permeable to ionic current-carrying species in a nonspecific manner.

8. An electrode as recited in Claim 7, wherein:

said reference membrane is formed of a hydrophobic polymeric material.

9. An electrode as recited in Claim 8, wherein:

said polymeric material is selected from the group consisting of cellophane, collodion, and cellulosic derivatives such as cellulose acetate butyrate, cellulose

acetate, cyanoethylated cellulose, cellulose propionate, cellulose tridecanoate, and ethyl cellulose.

10. An electrode as recited in Claim 7, wherein:

05 said polymeric material is applied to the outer surface of said second salt layer of said electrode by dipping of said electrode into a solution of said polymer in a nonaqueous solvent which will dissolve the polymer without dissolving said salt layers.

11. An electrode as recited in Claim 7, wherein:

 said reference membrane is formed of cellulose acetate butyrate.

12. An electrode as recited in Claim 7, wherein:

 said reference membrane is formed of ethyl cellulose.

13. An electrode as recited in Claim 7, wherein:

 said reference membrane has a thickness of 0.05 to 0.8mm.

14. An electrode as recited in Claim 1, wherein:

 said electrically conductive inner element is at least partially formed of silver;

05 said first salt layer is formed of AgCl; and
 said second salt layer is formed of KCl.

15. An electrode as recited in Claim 14,
wherein:
said electrode can be sterilized.

16. An electrode as recited in Claim 14,
wherein:
said electrode can be sterilized for medical
purposes by various techniques including the use of
05 radiation, gas, ultraviolet light, and steam.

17. An electrode as recited in Claim 15,
wherein:
said electrode can be sterilized using ethylene
oxide gas.

18. An electrode as recited in Claim 1, where-
in said electrode is an ion-selective electrode further
comprising:
an ion-selective membrane surrounding said
05 second salt layer, said membrane being selectively
permeable to a specific ionic species.

19. An electrode as recited in Claim 18,
wherein:
said ion-selective membrane is formed of a
plasticized hydrophobic polymeric material.

20. An electrode as recited in Claim 19,
wherein:
said polymeric material is selected from the
group of materials consisting of silicone rubber,
05 polyvinyl chloride, polystyrene, polyhydroxyethyl-
methacrylate, polyvinylidene chloride, and polyurethane.

21. An electrode as recited in Claim 18,
wherein:

05 said polymeric material is applied to the outer surface of said second salt layer by dipping of said second salt layer coating said first salt layer into a solution of said polymer in a nonaqueous solvent which will dissolve the polymer without dissolving said salt layers.

22. An electrode as recited in Claim 18,
wherein:

05 said ion-selective membrane includes a plasticizer selected from the group of materials consisting of phthalic acid derivatives, citric acid derivatives, adipic acid derivatives, and sebacic acid derivatives.

23. An electrode as recited in Claim 18,
wherein:

05 said ion-selective membrane is formed of a mixture of polyvinyl chloride, didecylphthalate, and a complexing agent selective to said specific ionic species.

24. An electrode as recited in Claim 23,
wherein:

said ion-selective membrane has a thickness of from 0.05 to 0.8mm.

25. An electrode as recited in Claim 18,
wherein:

05 said electrode is positioned in a lumen of a catheter, said catheter lumen having an orifice for exposing a portion of said electrode in juxtaposition with said orifice to the environment.

26. An electrode as recited in Claim 25, further comprising:

05 means for attaching said electrode to said lumen around said orifice and for sealing the inside of said lumen from the ambient environment.

27. An electrode as recited in Claim 26, wherein:

05 said means for attaching and sealing includes a second membrane formed from the same material as said ion-selective membrane, said second membrane formed such that said ion-selective membrane and said second membrane form a contiguous layer, said second membrane also forming a bond with said catheter to seal the inside of said lumen from the environment.

28. An electrode as recited in Claim 27, wherein:

said second membrane is formed of plasticized hydrophobic polymeric material.

29. An electrode as recited in Claim 28, wherein:

05 said polymeric material of said second membrane is selected from the group of material consisting of silicone rubber, polyvinyl chloride, polystyrene, polyhydroxyethylmethacrylate, polyvinylidene chloride, and polyurethane.

30. An electrode as recited in Claim 28, wherein:

05 said second membrane is formed of a mixture of polyvinyl chloride, didecylphthalate, and a complexing agent selective to said specific ionic species.

31. An electrode as recited in Claim 30,
wherein:

said electrically conductive inner element is
at least partially formed of silver;

05 said first salt layer is formed of AgCl; and
said second salt layer is formed of KCl.

32. An electrode as recited in Claim 31,
wherein:

said electrode can be sterilized.

33. An electrode for medical applications,
comprising:

an electrically conductive inner element;

05 a first salt layer on the surface of said elec-
trically conductive layer having a cation in common with
some portion of said electrically conductive inner ele-
ment, also having an anion;

10 a second salt layer coating said first salt
layer, said second salt layer having an anion in common
with said first salt layer, said second salt layer being
totally anhydrous; and

a catheter surrounding a portion of said inner
element and said first and second salt layers.

34. An electrode as recited in Claim 33,
wherein said electrode is a reference electrode further
comprising:

05 a first membrane surrounding said second salt,
said membrane being permeable to ionic current-carrying
species in a nonspecific manner.

35. An electrode as recited in Claim 34,
wherein:

05 said electrode is positioned in a lumen of
said catheter, said catheter lumen having an orifice for
exposing a portion of said electrode in juxtaposition
with said orifice to the ambient environment.

36. An electrode as recited in Claim 35, further comprising:

05 means for attaching said electrode to said
lumen around said orifice and for sealing the inside of
said lumen from the environment.

37. An electrode as recited in Claim 36, wherein:

05 said means for attaching and sealing includes
a second membrane formed from the same material as said
first membrane, said second membrane formed such that
said first membrane and said second membrane form a contiguous layer, said second membrane also forming a bond with said catheter to seal the inside of said lumen from the environment.

38. An electrode as recited in Claim 37, wherein:

 said membrane is formed from a hydrophobic polymeric material.

39. An electrode as recited in Claim 38, wherein:

05 said polymeric material is taken from the group consisting of cellophane, collodion, or cellulosic derivatives such as cellulose acetate butyrate, cellulose acetate, cyanoethylated cellulose, cellulose propionate, cellulose tridecanoate, or ethyl cellulose.

40. An electrode as recited in Claim 38,
wherein:
said membrane is formed of cellulose acetate
butyrate.

41. An electrode as recited in Claim 38,
wherein:
said reference membrane is formed of ethyl
cellulose.

42. An electrode as recited in Claim 35,
wherein:
said electrically conductive inner element is
at least partially formed from silver;
05 said first salt layer is formed from AgCl; and
said second salt layer is formed from KCl.

43. An electrode as recited in Claim 42,
wherein:
said electrode can be sterilized.

44. A sensing device for selectively monitor-
ing ionic activity, comprising:
a catheter having one or more lumens;
a first electrode disposed within a lumen of
05 said catheter, said lumen being provided with a first
orifice to expose a portion of said first electrode to
the environment, said first electrode being a nonselec-
tive reference electrode; and

a second electrode disposed within a lumen of
10 said catheter, said lumen being provided with a second
orifice to expose a portion of said second electrode to

the environment, said second electrode being selectively responsive to changes in a particular ionic component being monitored in said environment.

45. A device as recited in Claim 44, wherein:
said first and second electrodes are in the same lumen of said catheter, said electrodes being electrically insulated from each other.

46. A device as recited in Claim 44, wherein:
said first and second electrodes are in separate lumens of said catheter.

47. Any and all features of novelty described, referred to, exemplified or shown.

FIG. 1

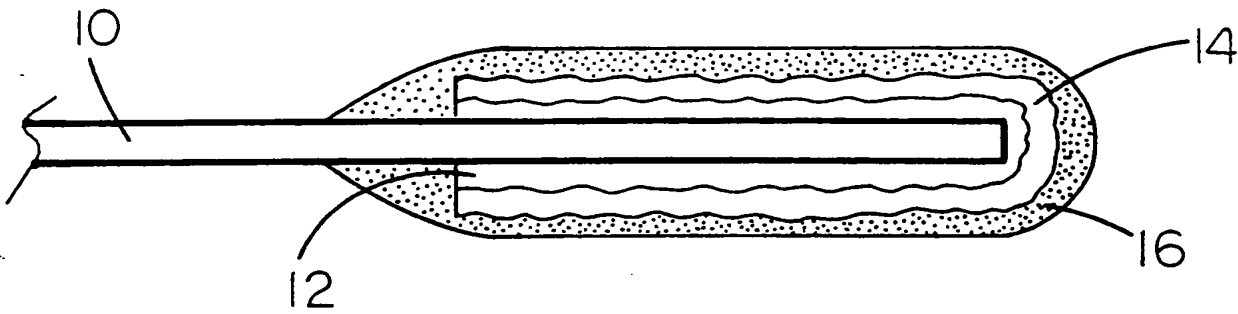


FIG. 2

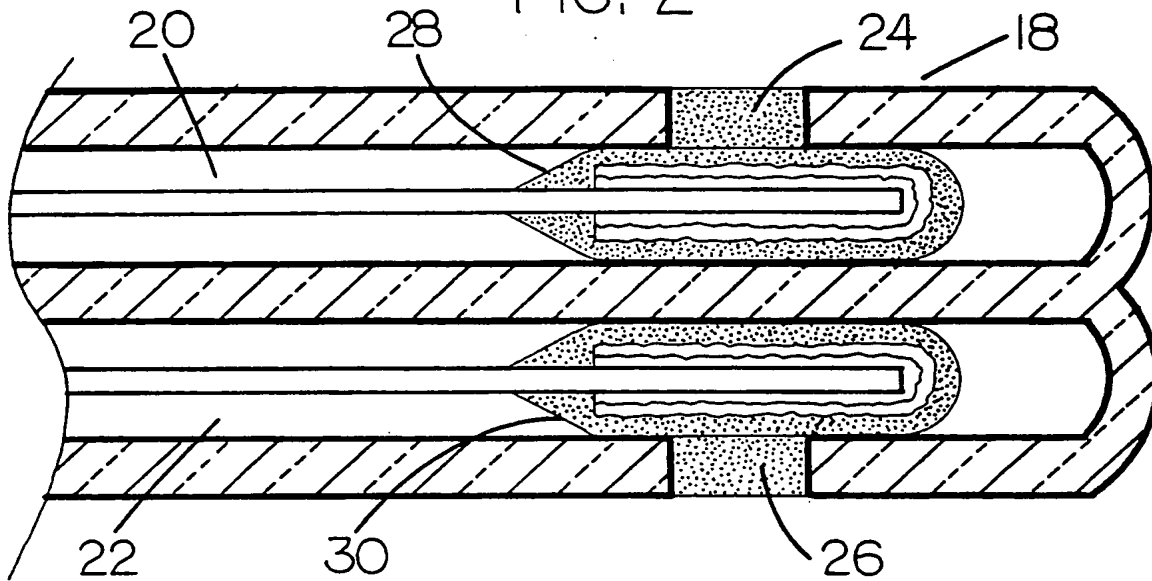


FIG. 3

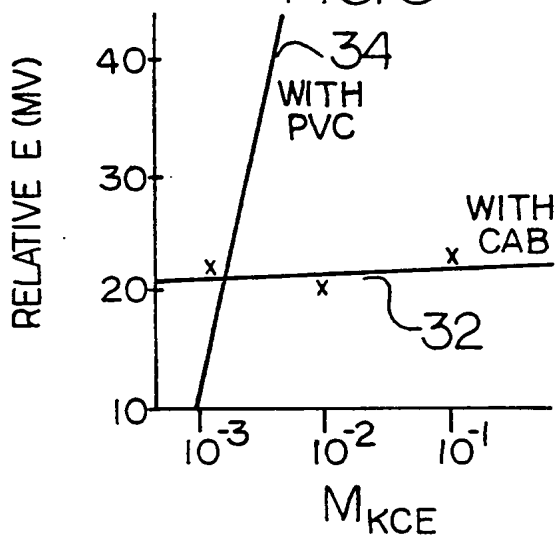
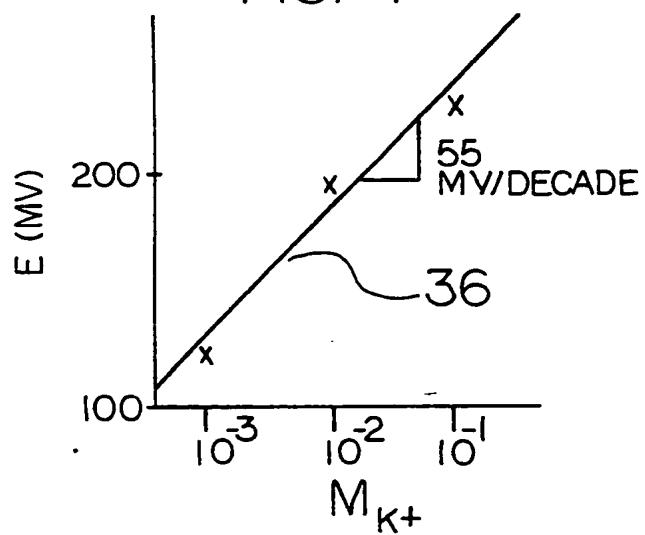


FIG. 4





European Patent
Office

EUROPEAN SEARCH REPORT

0193676

Application number

EP 85 30 7947

DOCUMENTS CONSIDERED TO BE RELEVANT

Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	CLASSIFICATION OF THE APPLICATION (Int. Cl.4)
Y	RESEARCH DISCLOSURE, no. 161, September 1977, page 29-39, Havant, GB; "Ion selective electrode" * page 29, column 2, last paragraph - page 30, column 1, paragraphs 1-3; page 30, column 1, "Metal/Metal-Salt Electrodes", paragraphs 2,4; page 30, column 2, paragraphs 3, 4 *	1,2,7, 8,10, 14,15, 18,19, 33,34, 44	G 01 N 27/30
Y	EP-A-0 074 198 (EASTMAN KODAK CO.) * page 6, lines 1-24; page 7, line 34 - page 8, line 4; page 17, lines 1-3, 18-29; claim 10 *	1,2,7, 8,10, 14,15, 18,19, 33,34, 44	TECHNICAL FIELDS SEARCHED (Int. Cl.4) G 01 N 27/00
Y,P	EP-A-0 155 638 (EASTMAN KODAK CO.) * page 1, lines 2-5, 10-11; page 3, lines 29-33; page 4, lines 1-4, 13-16; page 5, lines 7-14, 16-19, 24-26; claims * --- -/-	1,2,7, 8,10, 14,15, 18-21, 25,26, 33,34, 44,45	
The present search report has been drawn up for all claims			
Place of search BERLIN		Date of completion of the search 07-04-1986	Examiner VINSOME R M

CATEGORY OF CITED DOCUMENTS

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0193676

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Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	CLASSIFICATION OF THE APPLICATION (Int. Cl.4)
Y	WO-A-8 102 218 (J. KATER) * page 2, lines 5-22; page 5, lines 9-15; page 7, lines 21-23; page 14, lines 1-20; page 15, lines 26-30; page 16, lines 21-34; page 19, lines 13-22; page 20, lines 3-5; page 20, lines 13-18; claims *	1, 2, 7, 8, 10, 14, 15, 18-21, 25, 26, 33, 34, 44, 45	TECHNICAL FIELDS SEARCHED (Int. Cl.4)
Y	DE-A-1 922 225 (CORNING GLASS WORKS) * page 5, lines 13-24; page 7, lines 4-8; page 7, line 21 - page 9, line 15 *	1	
Y	GB-A-2 060 896 (RADIOMETER A/S) * abstract; page 2, line 65 - page 3, line 5; page 3, lines 28-31; claims *	1	
X	DE-A-2 722 617 (EASTMAN KODAK CO.) * claim 7; page 21, lines 13-18; page 23, lines 16-30; page 29, line 9 - page 30, line 2; page 30, lines 9-12, 17; page 32, lines 2-13, 24-30 *	1	
The present search report has been drawn up for all claims			

Place of search
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Examiner
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A	* whole document * ---	2-47	
A,D	US-A-3 856 649 (M.A. GENSHAW et al.) * whole document * --- -----	1-47	
The present search report has been drawn up for all claims			TECHNICAL FIELDS SEARCHED (Int. Cl.4)
Place of search BERLIN		Date of completion of the search 07-04-1986	Examiner VINSOME R M

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